

SUMMARY OF NOMENCLATURE

Formula starts with:	It must be:	Instructions
non-metal , but not NH ₄ or H or end in -COOH	molecular	use <u>prefixes</u> or it was a <u>memory</u> one (trivial) For states - molecular oxides are gases
a metal or NH ₄	ionic	name the positive ion, name the negative ion (no prefixes) For states - always solids unless water or (aq) is in reaction, then check solubility table
H or ends in COOH and has an (aq)	an acid	name it as hydrogen ----- then if it is: hydrogen -----ide, it becomes hydro----ic acid hydrogen -----ate, it becomes ----ic acid hydrogen -----ite, it becomes -----ous acid For states, always (aq)
Does the name:	It must be:	Instructions:
come from your memorized list?	molecular	write the memorized formula
use prefixes on both parts (but remember mono may be left off the first part) and is made of non-metals?	molecular	use the prefixes to tell you how many of each element
start with a metal or ammonium ion?	ionic	write the symbols and charges for the positive ion and the negative ion, then criss-cross the numbers on the charges and put in lowest terms
is ----- acid?	an acid	----- acid, follow these rules: If it is hydro----ic acid, it was hydrogen -----ide, balance charges If it is -----ic acid, it was hydrogen -----ate, balance charges If it is -----ous acid, it was hydrogen -----ite, balance charges to balance the charges for each ion, remember hydrogen is H⁺